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Naming binary compounds worksheet type 2

Type II Binary ionic compounds contain transition metals (including Group III, IV, V, VI metals, with the exception of Al) with non-metal ions. Display the correct name for the following connections. Provide correct formulas for these Type II binary compounds copper (II) iodine CuI_2 copper (I) oxide Cu_2O iron (II) sulfide FeS manganese (IV) oxide MnO_2 gold (III) chloride AuCl_3 gold (I) sulfide Au_2S silver bromide * AgBr cadmium sulfide* CdS vanadium (V) nitride V_3N_5 chromium(VI) oxide CrO_3 lead(IV) sulfide PbS_2 zinc selenide* ZnSe gallium(I) iodide GaI vanadium(III) chloride VCl_3 nickel(III) selenide Ni_2Se_3 iron (III) bromide FeBr_3 silver sulfide* Ag_2S tin (IV) chloride SnCl_4 manganese (II) fluoride MnF_2 antimony (V) bromide SbBr_5 cobalt(III) oxide Co_2O_3 titanium (IV) oxide TiO_2 Show correct Type II name for the following compounds FeCl_3 iron (III) chloride Au_2O gold (I) oxide FeCl_2 iron (II) chloride AuCl_3 gold (III) chloride V_2O_5 vanadium (V) 2 nitrogen PbO lead(II) oxide MnO_2 Crymenuse(IV) oxide Pb_2S_3 lead(III) sulfide NiF_3 nickel (III) fluoride Hg_2F_2 mercury(I) fluoride CuS copper(II) sulfide HgI_2 mercury(II) iodine CuBr copper (I) bromide CdSe cadmium selenid Fe_2O_3 iron(III) oxide ZnBr_2 zinc bromide* SnF_4 tin (IV) fluoride ScCl_3 scandium(III) chloride Ga_2S gallium(I) sulfid MnCl_2 crypts(II) chloride AgI silver iodide* NiCl_2 nickel(II) chloride *Please note that the transition metals zinc (Zn), cadmium (Cd) and silver (Ag) do not have multiple charges. Therefore, it is not necessary to place parantes after their names in connections, since they have fixed costs of +2, +2 and +1 respectively. Go Back to: Exam / Quiz Review Page | Composite naming page