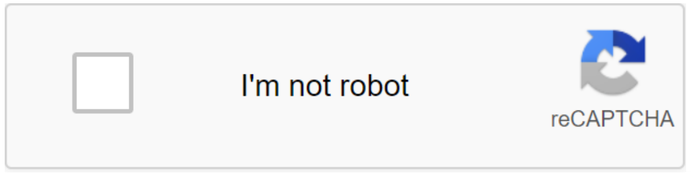

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Limiting reactants worksheet 12- 3 answer key

The opinions expressed by the participants of the entrepreneurs are their own. When I started my first company, at the age of 19, I had a special phone next to my bed to call clients. He often started ringing before dawn. Many of our customers were in other time zones. The second time, with my company Canva, I sleep until the alarm rings, but customer service is still at the heart of what we do. We built the company from scratch on excellent customer service. Related: The magic of owning customer service customer service problems is an important part of our company and every company. Here are the key things I learned: 1. Be where your customers are. Social media has changed the way companies work. More people now expect to be able to interact with your brand online. Your content and customer service strategy goes hand in hand. Users often ask questions about the product on social media. Make sure someone is available, around the clock, to answer questions. This is important for how users feel about your company. The deeper your understanding of how social media works, the better. 2. Keep your customers informed. According to a Kissmetrics survey, 81% of people are more likely to let a company repeat a business after a positive customer service experience. Keeping users up to date is the first step to avoid negative customer service. Related: 8 Keys to Customer Service Launch Awards When a customer contacts you with a problem, they hope for a solution. Go over and over again, keep them informed, and stay ahead if you can't immediately solve the problem. Make sure they have a positive experience. Most customers say they would prefer to access self-help information if it is available. Providing a support center is a great way to do that. 3. Give the customer a voice. Amazon founder and CEO Jeff Bezos has an empty chair at all his meetings. This provides a constant reminder that the most important person, the client, is not in the room. Keep in mind this invisible client when you make decisions. A customer-focused organization puts the customer first. Excellent customer service requires effort from your entire organization. Integrate your support team with marketing, product development, engineering, etc. Everyone on your team should understand the importance of happy customers. Regularly share customer feedback with the entire team. Sharing positive feedback helps boost the morale of the team and keeps everyone focused on your users. Customers is the backbone of every great company. It affects how you design your product, how you interact with customers, and how your team works. Related: Why stellar customer service is the key to building yours Brand responders needed for photosynthesis are carbon dioxide and water; they are obtained by combining pores on leaves and root systems in the ground. They eventually, through a process, turned into oxygen and glucose. The actual equation for photosynthesis is carbon dioxide plus water equals (in the presence of the sun) a certain type of glucose and oxygen. It takes six carbon dioxide molecules and six water molecules to go through a chemical process that, when added along with sunlight, converts them into an equation: C6H12O6-6O2. There are two separate reactions for this chemical change. Light-dependent reactions when sunlight breaks down water molecules apart (H2O is broken down into two hydrogen molecules and one oxygen) and the carrier molecules store hydrogen away for future use to make glucose for the plant. Solar energy is also used to change adenosine-dipphosphate to adenosine triphosphate (ADP to ATP). At night, when the plant lights independent, when the plant begins to take the molecules of the carrier and convert the hydrogen together with ATP into glucose. Chemical glucose formula C6H12O6. This shows that the plant used hydrogen (from the media molecules in which it was stored during the day), carbon dioxide (absorption during the day and night contains carbon dioxide in the air) and oxygen molecules to make its glucose to be stored for later energy. Limiting reagent or limiting reagent is a reagent in a chemical reaction that determines the amount of product that is formed. Identification of the limiting reactionary allows you to calculate the theoretical yield of the reaction. The reason why there is a limiting reactionary factor is that the elements and compounds react according to the ratio of moles between them in a balanced chemical equation. For example, if the ratio of moles in a balanced equation states that it takes 1 mole of each reaction to produce a product (1:1 ratio) and one of the reactionants is present in greater quantities than the other, the response is present in smaller amounts will limit the reaction. All this will be used before the other reactionary runs out. Given 1 mole of hydrogen and 1 mole of oxygen in reaction: 2 H2 and O2 No 2 H2O Limiting reactionary will be hydrogen because the reaction uses hydrogen twice as fast as oxygen. There are two methods used to find a limiting reaction. First, compare the actual ratio of moles to the ratio of moles to a balanced chemical equation. Another method is to calculate the gram mass of the product as a result of each react. Reacting, which gives the slightest mass of the product is limiting reactionary. Using the mole ratio: Balance the equation for chemical reaction. When converting masses of reactionary media into moles. If the number of reactionians is given in moles, skip this step. Calculate the mole mole between reactionaries using actual numbers. Compare this ratio with the mole-to-reaction ratio in a balanced equation. Once you determine which reactor is limiting reactionary, calculate how much product it can make. You can check that you have chosen the right reagent as a limiting reagent by calculating how much product the entire volume of the other reagent will give (which should be a larger number). You can use the difference between moles not limiting the reaction that is consumed and the starting number of moles to find the amount of excess reaction. If necessary, convert the moles back into grams. Using the product approach: Balance the chemical reaction. Convert this amount of reactionary means into moles. Use the mole ratio from a balanced equation to find the number of product moles that would be formed by each reaction if the full amount was used. In other words, perform two calculations to find the product's moles. Reacting, which gave a smaller amount of product is limiting reactionary. The reaction that gave more produce is the excess reactionary. The amount of excess reaction can be calculated by subtracting the moles of excess reaction from the number of moles used (or by subtracting the mass of excess reaction from the total mass used). Mole to gram unit conversions may be necessary to get answers to home problems. Reactionary means are the starting point in a chemical reaction. Reactions undergo chemical changes in which chemical bonds break down and new ones are formed to produce products. In the chemical equation, the reactionions are listed on the left side of the arrow, while the products are on the right side. If the chemical reaction has an arrow that points both left and right, then the substances on both sides of the arrow are reactionary as well as products (the reaction continues in both directions at the same time). In a balanced chemical equation, the number of atoms of each element is the same for reactionaries and products. The term reactionary first came into force around 1900-1920. The term reagent is sometimes used interchangeably General reaction can be given by equation: A q B qC In this example, A and B are reagents and C is a product. There shouldn't be a few reactions in reaction, however. In reaction decomposition, for example: C and B C is reactionary, while A and B are products. You can tell the reactionions because they are in the tail of the arrow that points to the products. H2 (hydrogen gas) and O2 (oxygen gas) are the reaction that forms liquid water: 2 H2 (g) and O2 (g) 2 H2O (l). There's a lot of mass in this equation There are four hydrogen atoms in both the reactionary and the product side of the equation and two oxygen atoms. The state of matter (s - hard, l - liquid, g - gas, aq - aqueous) is indicated after each chemical substance A balanced chemical equation shows the number of molar reactions that will react together to produce a molty number of products. In the real world, reactionary people rarely combine with the exact amount needed. One of them will be fully used before the others. Reactionary is used first known as limiting reactionary. Other reactionary products are partially consumed when the remaining amount is considered redundant. This example of the problem demonstrates a method for determining the limiting reactionary chemical reaction. Sodium hydroxide (NaOH) reacts with phosphoric acid (H3PO4) to form sodium phosphate (Na3PO4) and water (H2O) reaction: 3 NaOH (aq) 3 H2O (l) If 35.60 grams of NaOH reacts with 30.80 grams H3PO4, a. How many grams of Na3PO4 are formed?b. What is limiting to react?c. How many grams of excess reaction remains when the reaction is complete? Useful information: NaOH Molar mass - 40.00 gramsMolyar mass H3PO4 - 98.00 gramsMolar mass Na3PO4 - 163.94 grams To determine the limiting reaction, calculate the amount of product formed by each reaction factor. Reacting produces the least amount of product is limiting reactionary. To determine the number of grams Na3PO4 is formed: na3PO4 grams (gram reactionary) x (mole reaction-molar mass reactionion) x (mole ratio: product / reactionary) x (molar mass product / mole product) Number na3PO4 is formed from 35.60 4o gram NaOH gram Na3PO4 (35.60 g NaOH) x (1 ma NaOH/40.00 g NaOH) x (1 ma Na3PO4/3 mol NaOH) x (163.94 g Na3PO4/1 mol Na3PO 4) grams Na3PO4 48.64 grams Number Na3PO4 formed from 30.80 grams of H3PO4 grams Na3PO4 (30.80 g H3PO4) x (1 mall H3PO4/98.00 grams H3PO4) x (1 ma Na3PO4/1 mol H3PO4) x (163.94 g Na3PO4/1 mol Na3PO4)gram Na3PO4 - 51.52 g sodium hydrodoxide forms less product than phosphoric acid. This means that sodium hydroxide was a limiting reagent and formed 48.64 grams of sodium phosphate. To determine the amount of excess reaction remaining, you need the amount used. reactionary grams used (product gram formed) x (1 product mae/product mass mole) x (reaction/product maul ratio) x (reaction mass) H3PO4 used163.94 g Na3PO4) x (reaction mass) 1 mole H3PO4/1 mol Na3PO4) x (98 g H3PO4/1 mol) grams of used H3PO4 and 29.08 g This number can be used to determine the remaining amount of overreaction. H3PO4 remaining grams - initial grams of H3PO4 - grams H3PO4, remaining 30.80 grams - 29.08 grams H3PO4 remaining 1.72 grams When 35.60 grams NaOH reacts with 30.80 grams H3PO4, a. 48.64 grams Na3PO4 formed.b. NaOH was limiting reactant.c. 1.72 grams of H3PO4 remain in the completion stage. Complete. limiting reactants worksheet 12-3 answer key