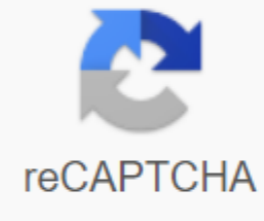




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Stoichiometry chapter 12 test a answers

1. Instant speed is the reaction rate at any given time, a period of time that is so short that the concentrations of reactionary and products vary by a small amount. The initial speed of the instant reaction rate as it begins (as the product is just beginning to form). The average rate is the average instant bet over a period of time. 3. Bet 12 ClF3 12'12'0'12'0'12'ClF3't'13'13'F2't 5. (a) Average speed, 0 - 10 s, 0.0375 mol L⁻¹ s⁻¹; average speed, 10 x 20 s and 0.0265 mol L⁻¹ s⁻¹; (b) instant speed, 15 s and 0.023 mol L⁻¹ s⁻¹; (c) Average formation speed B 0.0188 mol l/l s/l; Instant speed for the formation of B 0.012 mol l/l s/l. Higher molarity increases the speed of reaction. Higher temperature increases the speed of reaction. Smaller pieces of magnesium metal will react faster than larger pieces because more reactive surface exists. 9. (a) Depending on the angle chosen, the atom may take a long time to collide with the molecule, and when a collision occurs, it cannot break the connection and form another. (b) Reaction particles must come into contact with each other before they react. 11. (a) very slowly; (b) As temperatures rise, the response is faster. The number of reactionary means decreases and the number of products increases. After a while, there is approximately the same amount of B, C, AB, and C in the mix and a slight excess of A. 15. (a) The process reduces the rate by 4 per cent. (b) Because CO does not appear in the betting law, the rate is not affected. 19. 21. q k bet; 2.0 x 10 x 10 x 1 mol l/l h-1 (about 0.9 g L⁻¹ h-1 for the average male); The reaction is zero order. 23. NoCl-2 bet; 8.0 and 10 x8 l/mol/h; Second order 25. RATE no 2 'Cl2'; 9.1 litres of molh-2 h/1; Second Order in NO; The First Order in Cl2 27. (a) The Betting Act is the second order in A and is written as a rate of zak2. (b) to 7.88 x 10-3 l mole-1 with 129. (a) 2.5 x 10-4 ma/l/min 31. Rate of th CZIOKL; 6.1 x 10 x2 l mole No.1 with 133. The In SO2Cl2 graph reveals a linear trend compared to t; so we know it's a first-order

reaction: to 2.20 and 105 to 134. The plot is beautifully linear, so the reaction of the second order. 50.1 litres of mole-1 h-142. Reaction to the first order. 1.0×10^7 l mole⁻¹ min.⁻¹144. 1.16 and 103 s; 20% remains48. 0 (M) to 103 (c-1) 4.88 2.45 3.52 2.51 2.29 2.53 1.1. 81 2,58 5,33 2,36 4,05 2,47 2,95 2,48 1,72 2,43 50. Reactionary either can move too slowly to have enough kinetic energy to exceed the energy activation for the reaction, or the orientation of the molecules when they collide can prevent the reaction. 52. Activation energy is a minimum amount of energy to form an activated complex in reaction. It is usually expressed as the energy needed to form a single mole of the activated complex. 54. After finding k at several different temperatures, the $\ln k$ versus $1/T$ area gives a straight line to the E_a/R slope from which E_a can be defined. (a) 4 times faster (b) 128 times faster than 58. 64. E_a No 108 kJ; A 2.0×10^8 with 1; 3.2×10^{-10} x1; (b) 1.81×10^8 hours or 7.6×10^6 days; (c) Assuming that the reaction is irreversible, it simplifies the calculation because we do not have to take into account any reactor that, having been converted into a product, returns to its original state66. The atom has enough energy to react from BC; however, the different angles at which it bounces from B.C., without reacting, indicate that the orientation of the molecule is an important part of the kinetics reaction and determines whether the reaction will occur. 68. No. Overall, for the overall reaction, we cannot predict the effect of a change in concentration without knowing the law on rates. Yes. If the reaction is an elementary reaction, the doubling of the concentration of A doubles the speed. 70. Rate and ZAKSB2; Rate 3 72. Rate1 and CSO3; Tariff2 and CSO3 HL; (c) Tariff3 and CPC CLOSO; Tariff2 and CSO3'NO; (e) Rate3 and to NO2'O 74. (a) Doubling (H2) doubles the speed. H2 must introduce a law on rates to the first authority. Doubling NO increases the bet by 4 times. NO should introduce a law on rates in the second power. (b) Rate no 'NO'2'H2; (c) to 5.0×10^3 moles-2 l-2 min.1; (d) 0.0050 mol/l; Step II is a speed-defining step. If Step i gives N2O2 enough, steps 1 and 2 combine to give $2NO-H2 \rightarrow H2O-N2O$. $2NO-H2 \rightarrow H2O-N2O$. This reaction is in line with the observed betting law. Combine steps 1 and 2 with step 3, which occurs on the assumption in a quick way to give the appropriate stoichiometry. 76. The overall way of the catalyst is to provide a mechanism by which response measures can be more willing to combine if they follow the path of a lower-energy response. The speed of both forward and reaction increases, which leads to faster balance. 78. (a) Chlorine atoms are a catalyst because they react in the second stage but are regenerated in the third stage. Thus, they are not used, which is typical for catalysts. (b) NO is the catalyst for the same reason as part of the a). 80. Reducing the energy of the transitional state indicates the effect of the catalyst. (a) B; (b) B82. The energy needed to transition from the original state to the transition state is (a) 10 kJ; (b) 10 kJ.84. Both charts describe two-step, exothermic reactions, but with different variations in enthalpy, suggesting that the diagrams depict two different common reactions. In order to continue our website, we ask you to confirm your identity as a person. Thank you so much for the Cooperation. Cooperation.

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