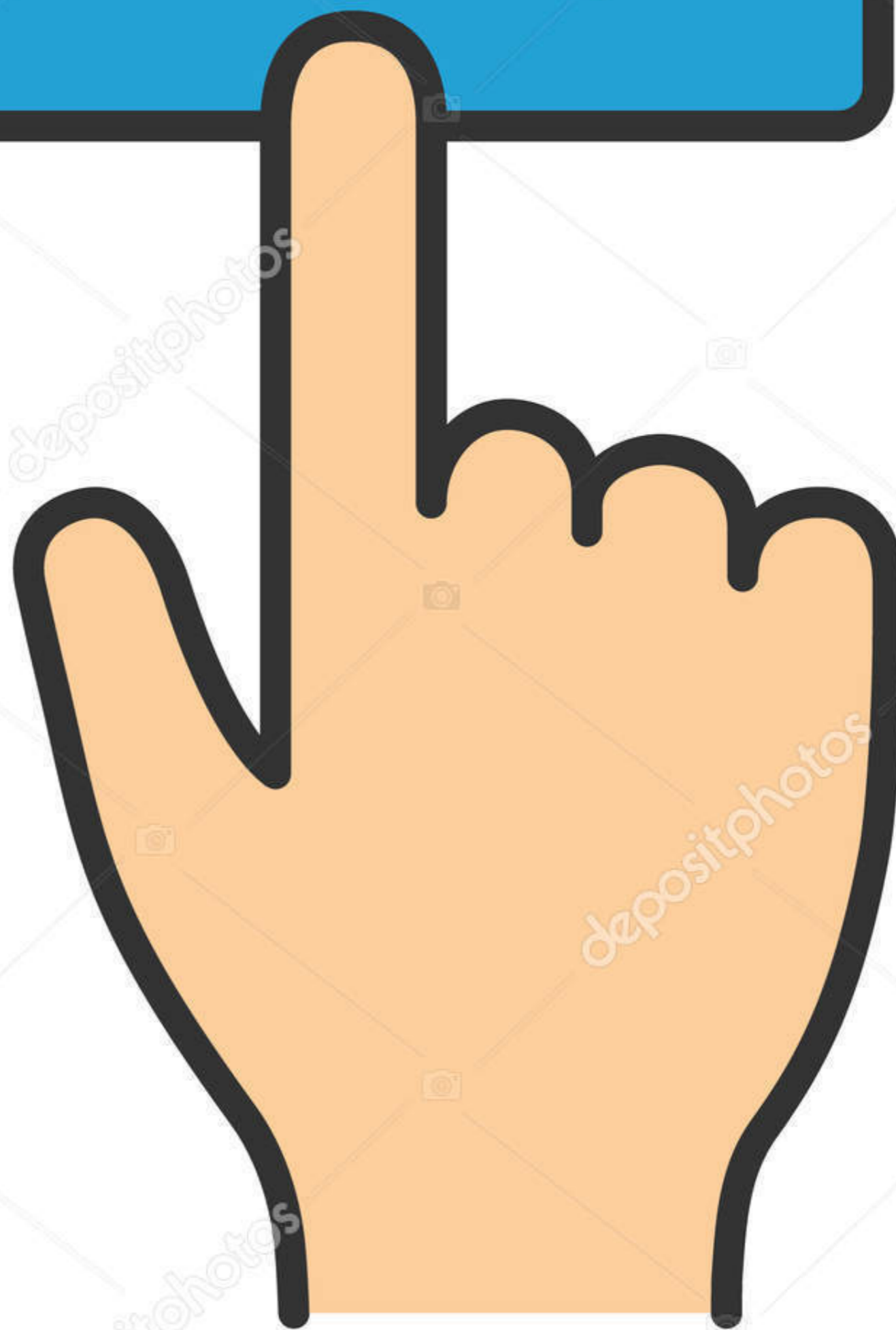
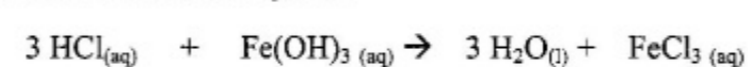


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## KEY

5. Given the balanced equation:



a) It takes 19.56 mL of 0.50 M HCl to titrate a 25.0 mL sample of a solution of  $\text{Fe}(\text{OH})_3$ . Calculate the  $[\text{Fe}(\text{OH})_3]$

$$\begin{aligned} \text{Moles HCl} &= 0.50 \text{ M} \times 0.01956 \text{ L} = 0.00978 \text{ mol HCl} \\ \text{Mole Fe}(\text{OH})_3 &= 0.00978 \text{ mol HCl} \times \frac{1 \text{ mol Fe}(\text{OH})_3}{3 \text{ mol HCl}} = 0.00326 \text{ mol Fe}(\text{OH})_3 \\ [\text{Fe}(\text{OH})_3] &= \frac{0.00326 \text{ mol Fe}(\text{OH})_3}{0.0250 \text{ L Fe}(\text{OH})_3} = \boxed{0.130 \text{ M}} \end{aligned}$$

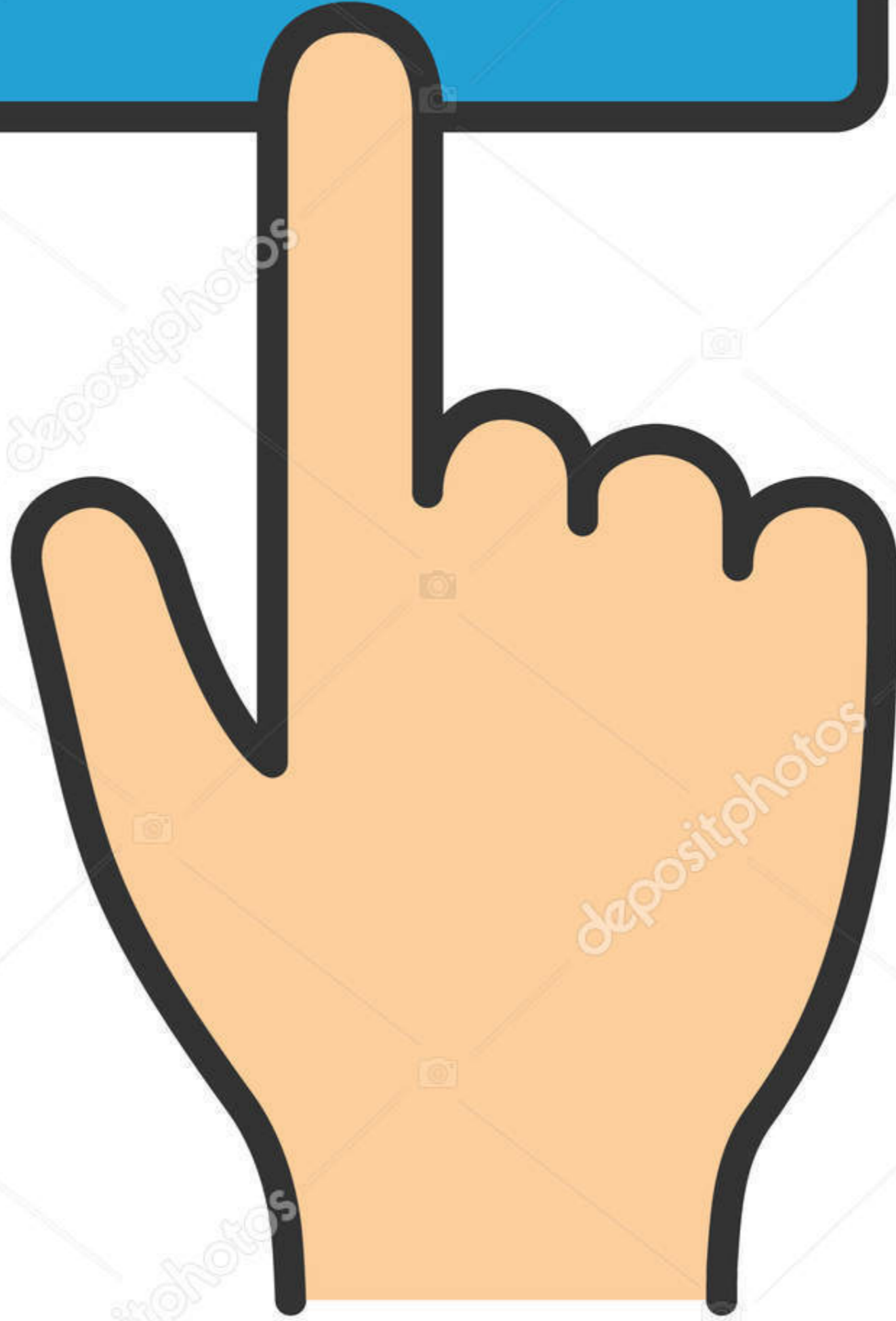
b) What mass of  $\text{Fe}(\text{OH})_3$  is needed to completely react with 10.0 mL of 0.50M HCl solution?

$$\begin{aligned} \text{moles HCl} &= 0.50 \text{ M} \times 0.0100 \text{ L} = 0.00500 \text{ mol HCl} \\ 0.00500 \text{ mol HCl} &\times \frac{1 \text{ mol Fe}(\text{OH})_3}{3 \text{ mol HCl}} \times \frac{106.8 \text{ g Fe}(\text{OH})_3}{1 \text{ mol Fe}(\text{OH})_3} \\ &= \boxed{0.178 \text{ g Fe}(\text{OH})_3} \end{aligned}$$

c) What volume of 0.50M HCl is required to titrate a 21.36 gram sample of iron (III) hydroxide?

$$\begin{aligned} 21.36 \text{ g Fe}(\text{OH})_3 &\times \frac{1 \text{ mol Fe}(\text{OH})_3}{106.8 \text{ g Fe}(\text{OH})_3} \times \frac{3 \text{ mol HCl}}{1 \text{ mol Fe}(\text{OH})_3} \\ &= 0.600 \text{ mol HCl} \\ \text{L} &= \frac{\text{mol}}{\text{M}} = \frac{0.600 \text{ mol}}{0.50 \text{ M}} = \boxed{1.20 \text{ L}} \end{aligned}$$

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