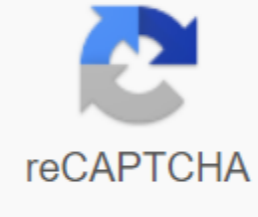




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## Chemistry chapter 3 mixed review answers

Transcript for Back Print Date Date Class 3 REVIEW Atoms: Building Blocks of Matter MIXED REVIEW SHORT ANSWER Answer Answer to the following questions in space provided. 1. The boron element, B, has an atomic mass of 10.81 amu according to the periodic table. However, no boron atom has a mass of exactly 10.81 amu. How can you explain this difference? 2. How do the results of Rutherford's gold foil experiment indicate the existence of the nucleus? 3. Ibuprofen, C<sub>13</sub>H<sub>18</sub>O<sub>2</sub>, which is produced in Michigan contains 75.69% mass carbon, 8.80% hydrogen and 15.51% oxygen. If you buy ibuprofen for a headache while you are on vacation in Germany, how do you know that it has the same percentage of makeup as ibuprofen you buy at home? 4. Complete the following chart using atomic mass values from the periodic table: Connection Mass Fe (g) Mass O (g) Ratio O:FeO Fe<sub>2</sub>O<sub>3</sub> Fe<sub>3</sub>O<sub>4</sub> MODERN CHEMISTRY Copyright © Holt, Rinehart and Winston. All rights are reserved. ATOMS: BUILDING BLOCKS OF MATTER 23 Back Print Name Date Date CLASS MIXED REVIEW continued 5. Fill the following table: Element Symbol Atomic Number Sodium Mass Number of Protons Number of Neutrons 22 F 9 19 80 40 45 20 1 0 222 PROBLEMS provided. Number of electrons 86 Write the answer on the line to the left. Show all your work in Space 6. How many atoms in 2.50 moles of hydrogen? B. How many atoms in 2.50 moles of uranium? C. How many moles are present in 107 grams of sodium? 7. A certain element exists as three natural isotopes, as shown in the table below. Isotopic Mass (Amu) Percentage of Natural Abundance Mass No. 1 19.99244 90.51 20 2 20.99395 0.27 21 3 21.99138 9.22 22 Calculate the average atomic mass of this element. 24 ATOMS: BUILDING BLOCKS OF MATTER MODERN CHEMISTRY Copyright © Holt, Rinehart and Winston. All rights are reserved. Back Print Title Date Class Chapter 3 REVIEW Atoms: Building Blocks of Matter MIXED REVIEW SHORT ANSWER Answer the following questions in space provided. 1. The boron element, B, has an atomic mass of 10.81 amu according to the periodic table. However, no boron atom has a mass of exactly 10.81 amu. How can you explain this difference? The periodic table shows the average atomic mass, which is weightedly for all boron isotopes. 2. How do the results of Rutherford's gold foil experiment indicate the existence of the nucleus? Several alpha particles have bounced off and therefore should have hit a dense bundle of matter. 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