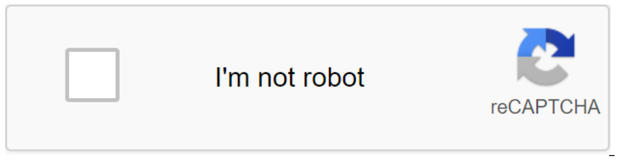

Continue

A level chemistry paper 5 notes pdf

Hello everyone, here's Chemistry Level Papers 5 solved documents in which the answers are written in full form as we know that CIE labeling schemes are sometimes hard to understand, plus it doesn't contain the appropriate diagrams. Syllabus Code: 9701 As we all know that Document 5 can be very difficult, especially when starting to answer a question, we do not know where and how to start especially question 1 where you have a plan for the whole experiment. So I attached completely settled documents that I wrote recently. Hope these resolved documents are useful 😄 Link to P5 solved documents are below the link: Chemistry Papers 5 (P5) Resolved past documents For more note and past paper, please click on the Home Thank You for Planning section: 1. When asked to draw a diagram, ALWAYS mention the amount of device used. The most commonly used is a gas syringe. I think it's best to limit the volume of the syringe to less than 500 cm3. In addition, the volume of the widely used small test tube is about 16 cm3, and the volume of boiling tube is about 25 cm3. We should always keep the amount of the machine that we use when choosing the scope of the planning solution. Saying that we are going to place 50cm3 aqueous HCl in a test tube is certainly not going to please the examiners. I think I ended up doing the same thing in one of the works where we had to prepare different concentrations of this solution. I chose the amount of water to be added to the glass to be 250 cm3. However, the ER stated that the volume should be less than 200cm3, which is the case since the volume of the widely used glass glass is about 250cm3. Someone holding my expm would have completely filled the glass. However, we can use more volume if we prepare that the volume of the glass is 400 cm3 or 500 cm3 or so on. So, about being the volume of the apparatus is very, very important 2. Always measuring the volume of solutions, use a burette or pipette because they have a low PERCENTAGE error. The question of percentage errors is usually asked in P5. Percentage error becomes very unacceptable if we measure small volumes of solution or small masses. Thus, the 3dp balance is much better for measuring masses than 2dp balance as it would be a much smaller % error when small masses are measured 3. When the crystals are heated, do not use a simple test tube or any machine with sharp edges, as they can crack on sharp edges. For a very strong heating to a constant mass fits the crucible placed on a tubular clay triangle. And of course, to measure the mass of crucible and solid, you will have to file it on the balance. But we wouldn't want to fry our bad balance by placing a very hot crucible on it. use magnesium tape or any dirty dirty Clean it with SANDPAPER How to prepare crystals (which is also associated with solubility) Many techniques are mentioned in MS, but I found that this is the simplest: CIE usually requires us to prepare a crystal after preparing a rich crystal solution. Thus, we must prepare a saturated solution first: 1. Take a fixed amount of water in a glass of appropriate volume 2. Add the crystal to the water and stir continuously. You have to give some time for the crystal to dissolve as it is the process of equilibrium 3. After 5 minutes of stirring, if solid crystals do not appear, add an additional mass of crystal 4. And repeat the process until the solids appear in the glass 5. The filter solution is using filter paper and funnel, so that the saturated solution is collected into the glass (whose mass was measured earlier) under the funnel Now we have a saturated solution in the glass. How to get crystals: 1. We have to place the glass in a warm bath with water. We could use a burner, but there is a risk of overheating of a solution that can decompose the crystal. If the heat is appropriate, the water solution should evaporate and we should have the dry crystals ready. How to measure solubility: 6. Measure the mass of the glass and solution from step 5 7. Subtract the mass of the glass from the mass to 6 to get a mass of saturated solution 8. And evaporate, as shown above, to get a mass of crystals. 8. Measure the mass of crystal glass 9. Subtract the mass of the glass from the mass of 8 to get a mass of crystals 10. Subtract the mass of crystals from the mass of the saturated solution to get a mass of water in solution 11. I assumed that all the masses are in grams. So to get solubility: Crystal mass x 100/Mass water When you have to remove moisture from : 1. Surface: Wash the surface with propanone flow. The water dissolves in the propanone and repeats it several times. Then gently heat the surface to evaporate the propanone from Surface 2. Par: Use dessicants such as: 1. ANHYDROUS sulphuric acid 2. CALCIUM CHLORIDE ANHYDROUS 3. Silica gel You have to pass steam from the glass, Containing dessicant Also, the useful property of lime soda is that it absorbs both water vapor and carbon dioxide We are also regularly asked to measure enthalpy changes Most of us already know that we are using a plastic cup and thermometer for this purpose however it has many drawbacks (asked regularly) and here are some with the necessary measures: 1. Heat loss in the neighborhood: Cover plastic cup with a lid. The air in the glass is a good insulator. C. Use a few cups to thicken the side layer of plastic 2. Instability cup a. Place the cup in a glass glass 3. For exothermic reactions, solution spray is very likely. Use a large glass for experiment, not a small plastic cup (which has a small volume) b. Put the lid on top of the glass (it's only MINIMISES spray, does not prevent it completely) 4. When we heat the volume of water in a glass glass, there are two cases of heat loss to consider: a. Loss of heat from a glass of water b. Loss of heat from heating the burner with the volume of water while preparing a solution of a fixed concentration from the given parent solution concentration, say, 2.0 ma/dm3 A. We are obliged to prepare a solution of concentration of 1.0 maul/dm3 with a volume of 250 cm3 (conn can not be more than 2 !!!!!!!!!!!!!!) a. Add 100 cm3 of parent solution to the bulk flask marking at 250 cm3 b. Top with water up to 250 cm3. Use a washing bottle for this purpose to have more control over the process of adding If necessary prepare 250 cm3 of 0.5 maul/dm3 crystal solution Mr 50g First understand that we only need 250 cm3, not 1 dm3 Now the 250 cm3 will be: 50/4 and 12.5 g solid So, first add 12.5 g of solid to 50 cm3 water in BEAKER (not bulk flask yet). It should also be noted that the volume of water is less than 250 cm3. Stir properly, and if the solid does not dissolve add more water until it dissolves completely then transfer the solution from the glass to a voluminous flask. Be sure to rinse the glass with water and transfer the solution in the tomtric flask Stopper flask and shake properly Finally, add the required amount of water to make the solution up to 250 cm3 USE A FUNNEL or YOU RISK OVERSHOOTING MARK Titration is accurate because: 1. The standard acid/base solution is used 2. we get concordant credits 3. % error in pipettes and buretts is very small 4. The end point of the credits is Sharp Credits for studyguide and Amna Shahid for exchange. PastPapers.Co Index/CIE-Revision/A-Level/A-As-Level-Chemistry-Notes-9701 Past documents by © 2020, PastPapers.Co Chemguide: Support CIE A level Chemistry Paper 5 Menu Reference topics for paper 5 Pages in this section cover a number of topics arising from my analysis of recent exam papers, mark schemes and examiner reports for paper 5. Introduction of Formula and Equation Calculations Graphics Tomtric Apparatus and Methods Drawing Chart Selection of Reasonable Parts of the Device Safety Considerations Go to CIE Chemistry Support Main Menu . . . Go to Chemguide Main Menu. . © Jim Clark 2017 For each of the papers below, there are revision notes, summary sheets, questions from past exam papers, divided by topic and other sheets. Micro961 Starter Thread Start Date May 15, 2014 2014 a level chemistry paper 5 notes pdf. z notes chemistry a level paper 5