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	valence electrons	electron dot structure showing bonds	# single bonds	# double bonds	# triple bonds	# unshared pairs	3-D drawing of molecule
H <sub>2</sub> O	8	H- $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{O}}}$ -H	2	0	0	2	
HCP	10	H-C $\equiv$ $\overset{\cdot\cdot}{\text{P}}$	1	0	1	1	
PH <sub>3</sub>	8	H- $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}}$ -H   H	3	0	0	1	
P <sub>2</sub> H <sub>2</sub>	12	H- $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}}$ - $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}}$ -H	2	1	0	2	
C <sub>2</sub> H <sub>2</sub>	10	H-C $\equiv$ C-H	2	0	1	0	
NCl <sub>3</sub>	26	$\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{N}}}$ - $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}}$ - $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}}$ - $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}}$	3	0	0	10	
H <sub>2</sub> S	8	H- $\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{S}}}$ -H	2	0	0	2	

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