In a sample consisting of 1.00 mol nabr and 0.300 mol ki, what is the mass percent of iodine?

Continue

Molarity Chemical Solutions Now mole solute are permanent. The volume of the solution can change significantly when the temperature rises or decreases. In some calculations, molyality is used in preference, which is determined by the factor.... this temperature is independent, and at lower concentrations, molarity - molness. To get molarity, you divide the moles of the solution into liters of solution. For example, a solution of 0.25 mole/l NAOH contains 0.25 moths/l naoH containing 0.25 moles of sodium hydroxide in each litre of solution. To calculate the molarity of the solution, it is necessary to know the number of moles of the solution and the total volume of the solution. To calculate molness: Calculate the number of moles soluble present. Calculate the number of litres present solutions. Divide the number of moles of the solution by the number of liters of solution. EXAMPLE: What is the molarity of a solution prepared by dissolving 15.0 g of NaOH in sufficient water to make a total of 225 ml of solution? Solution: 1 moth NaOH has a mass of 40.00 g, so The Moles NaOH and 15.0 cancel (g NaOH) × 1 ma naoh/ (40.00 cancel (g NaOH)) × - 0.375 Na molOH 1 L soln/ (1000 undo (mL soln)) - 0.225 L soln - Molarity - moles of solute/litres of solution - 0.375 mol/0.225 L - 1.67 mol/L Some students prefer to use the mole triangle. It summarizes the formulas of molarity, as moles - moles × liters - Mole/litres - Litres - moles/mole/molarity - Example #1: Calculate the amount of water (in kilograms), which should be added to 12.0 grams of urea, $(NH_2)_2CO$, when preparing 18.3 percent mass solution. The molar mass of urea $(NH_2)_2CO$ is 60,055 g/mole. g/mole.